



Short communication

Selenium functionalized carbon for high dispersion of platinum–ruthenium nanoparticles and its effect on the electrocatalytic oxidation of methanol

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H I G H L I G H T S

- Vulcan carbon can be modified by selenium atoms.
- Se on the surface of carbon affects the dispersion of PtRu particles.
- PtRu/Se–C showed higher electrocatalytic activity toward methanol oxidation.

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Using selenium functionalized carbon as supports, platinum–ruthenium nanoparticles were highly dispersed on the carbon surface, and showed improved electrochemical properties for methanol electrooxidation. The method provides a new route for functionalization of the carbon surface on which to disperse noble metal nanoparticles for application as electrocatalysts in fuel cells.

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1. Introduction

Platinum–ruthenium (PtRu) bimetallic electrocatalysts have been considered as promising electrocatalysts for methanol oxidation reaction (MOR) in direct methanol fuel cells (DMFCs) due to their high activities and CO tolerance [1]. PtRu catalysts are generally dispersed as nanoparticles (NPs) on a carbon support in order to obtain optimum catalyst utilization for methanol electrooxidation by increasing the dispersion of the nanosized catalysts. Therefore, highly distributed PtRu NPs with small size and narrow

size distribution are ideal for high electrocatalytic activity owing to their large surface-to-volume ratio [2]. It is well known that the physicochemical properties of the support can greatly influence the dispersion and interaction between catalyst NPs with the support, and consequently their electrochemical behavior. From this perspective, some of the recent studies have employed new carbon materials, such as, carbon capsules [3], graphitic carbon nitride [4] and C₆₀ [5] as supports to achieve a desired distribution of the PtRu NPs on the surface of the supports, and these catalysts exhibited improved performance for methanol oxidation. However, these new carbon materials do not yet meet the target for practical application because the complex synthesis methods are not suitable for industrial production.

Another promising route is the use of chemical functional groups, such as carboxylate [6], 1,6-hexamethylenediamine [7], 2-aminophenoxazin-3-one [8], and 3,4,9,10-perylene tetracarboxylic

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acid [9] to functionalize the surface properties of carbon materials to enhance the dispersion of PtRu NPs, thus increasing their utilization and electrocatalytic activity. In the last two decades, there has been increasing evidence to show that the electrochemical and physical properties of carbon materials are extremely sensitive to heteroatoms, such as N atoms, functionalizing the carbon surface or doping into the carbon structure [10–13]. Heteroatom functional groups introduced into carbon support appear to influence not only the active sites of the carbon support, which results in smaller catalyst particle size and increased catalyst particle dispersion, but also increases support/catalyst chemical binding (or “tethering”), which in turn results in enhanced stability of the catalysts. Furthermore, the metallic NPs electronic structure is modified, which enhances intrinsic catalytic activity [14]. In spite of these advantages, developing new heteroatom functionalities for modification of the carbon materials that provide well-dispersed noble metal NPs are still desirable.

Selenium, an element of the chalcogenide group and a semi-metal, has several advantages over the widely studied conventional heteroatom doping elements: 1) the electric conductivity of selenium facilitates electron transport during the electrochemical reactions [15]; 2) voltammetric studies show that selenium is stable in acid solution and shows a double layer region from -0.2 to 0.75 V (vs. SHE), which is wider than that of Pt (-0.4 to 0.7 V (vs. SHE)) [16]; 3) the catalyst may show better stability at high temperature due to the anchor effect of selenium atoms [17]; 4) Selenium can become metallic when bound to metals resulting in exceptional chemical stability in acidic media, even at high positive potential [18]. Thus, the synthesis of selenium functionalized carbon and a study of their support effect are of great interest.

Herein, by selecting carbon black Vulcan XC-72 as the support, because it is a commercial and widely used support in electrocatalysis, an alternative strategy for dispersing PtRu alloy NPs on carbon is reported. In detail, triphenyl phosphine selenide (PPh_3Se) resulting from triphenyl phosphine (PPh_3) and selenium (Se) was non-covalently coupled to the surface of the conventional oxidised carbon (as shown in Fig. 1). The resulting PPh_3Se and carboxylate modified carbon were heat treated to remove PPh_3 . Thus the surface of carbon can be modified by non-covalent adsorption of selenium. Using of selenium functionalized carbon (Se–C) as supports, highly dispersed PtRu NPs with narrow size distribution on carbon were synthesized and showed higher electrocatalytic activity and stability toward methanol electrooxidation than that supported on oxidized carbon (C).

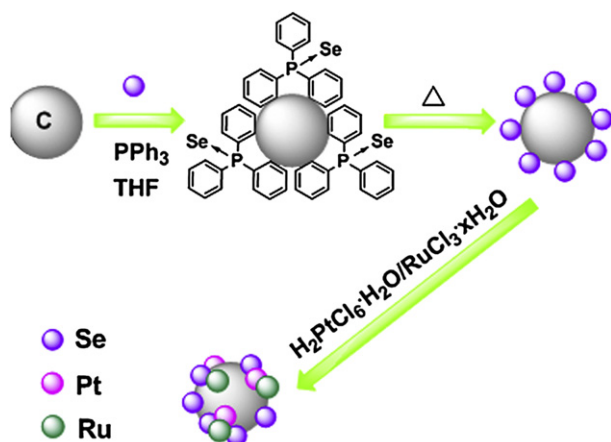


Fig. 1. Schematic diagram for Se-functionalized carbon and dispersion of PtRu nanoparticles.

2. Experimental

2.1. Synthesis

The procedure for selenium functionalization of carbon was prepared by a modified organic colloid method [19]. Carbon black (Vulcan XC-72R, Carbot Corp.) was pretreated by washing with acetone, oxidized with 2 mol mL^{-1} HNO_3 and 30% H_2O_2 , then washed with ultrapure water and dried overnight at 80°C , the pretreated carbon black (C) was obtained. 0.185 g of triphenyl phosphine were dissolved in 30 mL of tetrahydrofuran in a flask whilst stirring for 20 min , and 0.0504 g of elemental selenium was added and dissolved with stirring, then 1.002 g of the pretreated C were dispersed by ultrasonically into the solvent. The above solvent was transferred into a rotary dryer, and tetrahydrofuran was evaporated at 40°C . The mixture was dried in air at 60°C , and thermally treated in a tubular furnace under a N_2 atmosphere at 500°C for 2 h . The selenium functionalized carbon (Se–C) sample was obtained.

The detailed preparation process of the catalyst according to the literature [20] could be described as follows: 0.50 mL of RuCl_3 (20 mg mL^{-1}) solution and 0.97 mL of H_2PtCl_6 (20 mg mL^{-1}) solution were dissolved to a 100 mL flask with 30 mL ethylene glycol (EG) under stirring conditions. 500 mg of sodium citrate were dissolved the above solution, then 156 mg of sodium formate were added into the solution. pH of the system was adjusted to ~ 11 by adding 5 wt\% KOH/ethylene glycol solution dropwise with vigorous stirring. 100 mg of Se–C was added into the flask. The mixture was sonicated for 20 min and stirred 20 min to disperse the Se–C. The resulting solution was heated at 160°C for 8 h . After filtering and washing with ultrapure water, the sample (denoted as PtRu/Se–C) was obtained by dried overnight in air at 60°C for 24 h . For a comparison, the PtRu nanoparticles supported on the pretreated carbon black (PtRu/C) electrocatalysts were also prepared by using the same procedure.

2.2. Characterization

The obtained materials' structure was characterized by recording their X-ray diffraction (XRD) patterns on a Shimadzu XD-3A (Japan), using filtered Cu $K\alpha$ radiation (40 kV , 40 mA). X-ray photoelectron spectroscopy (XPS) was obtained by a PHI-5702 multifunctional X-ray photoelectron spectrometer (American). The average chemical compositions of PtRu/Se–C and PtRu/C catalyst were determined using an IRIS advantage inductively coupled plasma atomic emission spectroscopy (ICP-AES) system (Thermo, America) and EDX coupled TEM.

2.3. Electrochemical measurements

The electrochemical measurements of catalysts were performed using a CHI650 electrochemical work station. A conventional three-electrode electrochemical cell was used for the measurements, including a platinum wire as the counter electrode, an Ag/AgCl (3 M KCl) electrode as the reference electrode, and a modified glass carbon electrode (5 mm in diameter, 0.196 cm^2) as the working electrode was prepared as follows: 5 mg of catalyst was dispersed ultrasonically in 1 mL of Nafion/ethanol (0.25 wt\% Nafion). About $8 \mu\text{L}$ of the dispersion was transferred onto the glassy carbon (GC) disc using a pipette, and then dried in the air to the preparation of the catalyst layer on it. Before each measurement, the solution was purged with high-purity N_2 gas for at least 30 min to ensure the gas saturated. The recorded cyclic voltammograms are the 25th circles.

3. Results and discussion

Raman spectra of Se–C and C are shown in Fig. 2a peak at $\sim 1310 \text{ cm}^{-1}$ can be assigned to the A_{1g} breathing mode of the

disordered graphite structure (i.e., the D band), and the peak at $\sim 1585\text{ cm}^{-1}$ assigned to the E_{2g} structural mode of graphite (i.e., the G band). Comparison of the Raman spectra of Se–C and C shows that D band and G band do not significantly change in both width and shape after functionalization. The extent of the defects in graphite materials can be quantified by the intensity ratio of the D to G bands (i.e., I_D/I_G). The I_D/I_G ratio of Se–C (1.13) is smaller than that of C (1.18), suggesting that the sp^2 carbon network of graphene was not damaged by functionalization [21–23], so, the carbon support appears to retain its original structure. The slight decrease of the I_D/I_G ratio after selenium functionalization of the carbon may be attributed to the fact that the original defect sites on the carbon surface are covered by the selenium atoms, as reported in literature for the 3,4,9,10-perylene tetracarboxylic acid-functionalized CNTs [9].

The FT-IR spectra of C and Se–C are plotted in Fig. 2b. In the two curves, O–H band of OH group located at $\sim 3440\text{ cm}^{-1}$ and the C=O band of carboxylic acid groups at $\sim 1640\text{ cm}^{-1}$ are presented. These groups were introduced upon treatment with H_2O_2 and HNO_3 [24]. The antisymmetric C–Se stretching at 605 cm^{-1} , the symmetric C–Se stretching at 507 cm^{-1} and the C–Se bending mode at 462 and 436 cm^{-1} cannot be clearly assigned in the FT-IR spectrum of Se–C [25]. The results indicate possible non-covalent bonding of the selenium atoms. No characteristic diffraction peaks of PPh_3 are observed, implying that PPh_3 was removed by heat treatment.

The morphology of Se–C was characterized by transmission electron microscopy (TEM). Fig. 2c shows that Se–C particles are in powder form, which is consistent with other reports of Vulcan carbon morphology [26]. Further, it can be clearly observed that no heterogeneous NPs exist on the carbon surface, which was also confirmed from the high resolution transmission electron microscopy (HRTEM, the insert in Fig. 2b). Based on these analyses, it is conceivable that selenium atoms are uniformly attached onto the carbon surface by non-covalent adsorption.

Fig. 3a and b shows the TEM images of the PtRu/Se–C and PtRu/C samples. As shown in Fig. 3a, the Se–C is successfully decorated with many well-dispersed PtRu NPs. No evidence was observed of larger agglomerated particles on the Se–C surface. However, PtRu NPs aggregation clearly presents on the C surface. The distribution of PtRu NPs in the PtRu/Se–C catalyst shown as inset in Fig. 3a and

b is narrow compared to that of PtRu/C catalyst. This is due to the presence a large number of Se atoms on the C surface with uniform distribution that serve as the functional groups for the immobilization of PtRu NPs, leading to high dispersion and narrow distribution. EDX study of the PtRu/Se–C catalysts shown in Fig. 3c indicates the presence of Pt, Ru, Se, C and O, while that of the PtRu/C catalysts in Fig. 3d indicates the presence of Pt, Ru, C and O. The loading mass of PtRu and Se of the PtRu/Se–C catalysts determined by ICP is 9.86% and 3.83%, respectively, with 1:1 of Pt/Ru atom ratio. The loading mass of PtRu of the PtRu/C catalysts is 9.92% with 1:1 of Pt/Ru atom ratio, suggesting Pt and Ru are completely reduced. XRD was used to analyze the crystalline structure of the PtRu/Se–C and PtRu/C samples. No peaks from elemental Se were detected in the PtRu/Se–C sample (see Fig. 3e); thus, elemental Se appears to be homogeneously dispersed on the carbon particles [27], which is consistent with the TEM result. Before and after Se functionalization carbon, the crystalline structure of PtRu is not altered by the coexisting Se atoms. However, the electronic structure of Pt in the PtRu/Se–C sample clearly changes as shown by XPS (Fig. 3f). The Pt binding energies of the PtRu/Se–C and PtRu/C samples shift to the higher energy as compared to those of Pt/C, which can be attributed to the stronger metal-to-support interaction and a smaller particle size [11,28]. The Pt binding energies of the PtRu/Se–C shift to the lower energy by ca. 0.5 eV in related to those of PtRu/C samples, which is attributed that electron transfer occurs from Se to Pt (electronegativity: Pt = 2.28, Se = 2.40).

Fig. 4a shows the cyclic voltammograms (CVs) of the PtRu/Se–C and PtRu/C catalysts measured in a nitrogen-saturated 0.5 M H_2SO_4 solution. The two catalysts exhibit hydrogen adsorption/desorption peaks in the potential range from -0.20 to 0.10 V , and oxide formation/stripping wave peaks in the potential range from ca. 0.30 – 1.0 V . As can be seen from Fig. 4a, the onset potentials for the oxide reduction in the negative-going sweep are apparently shifted to more positive potential for the PtRu/Se–C catalysts as compared to PtRu/C. This shift indicates that the PtRu structure inhibits chemisorption the oxygenated species such as OH_{ad} chemisorb on the active sites of the PtRu/Se–C catalysts at high potential due to the change of the electronic structure induced by the presence of Se.

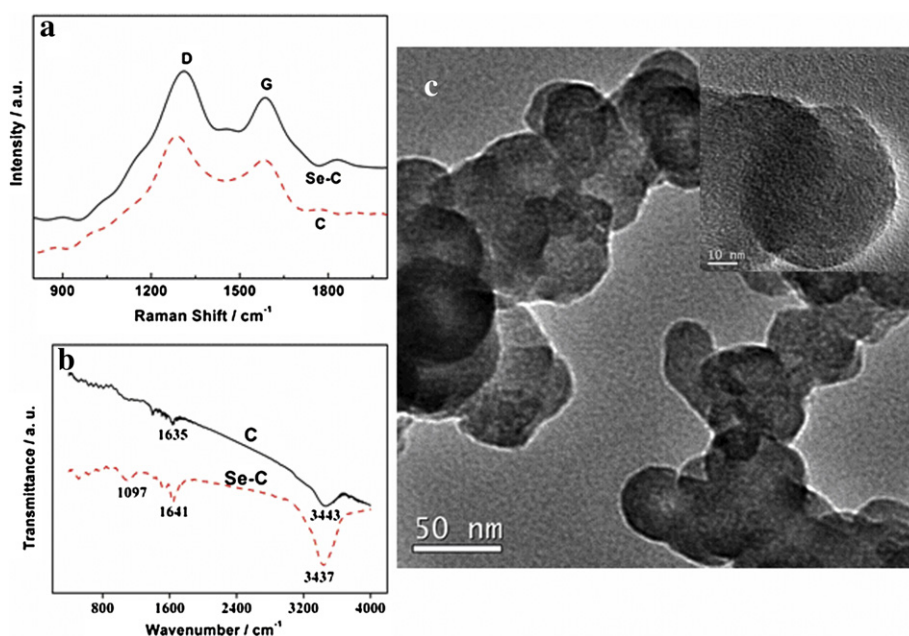


Fig. 2. Raman spectra (a) and FT-IR (b) of Se–C and C; The TEM image of Se–C (c), Inset of a HRTEM image of Se–C.

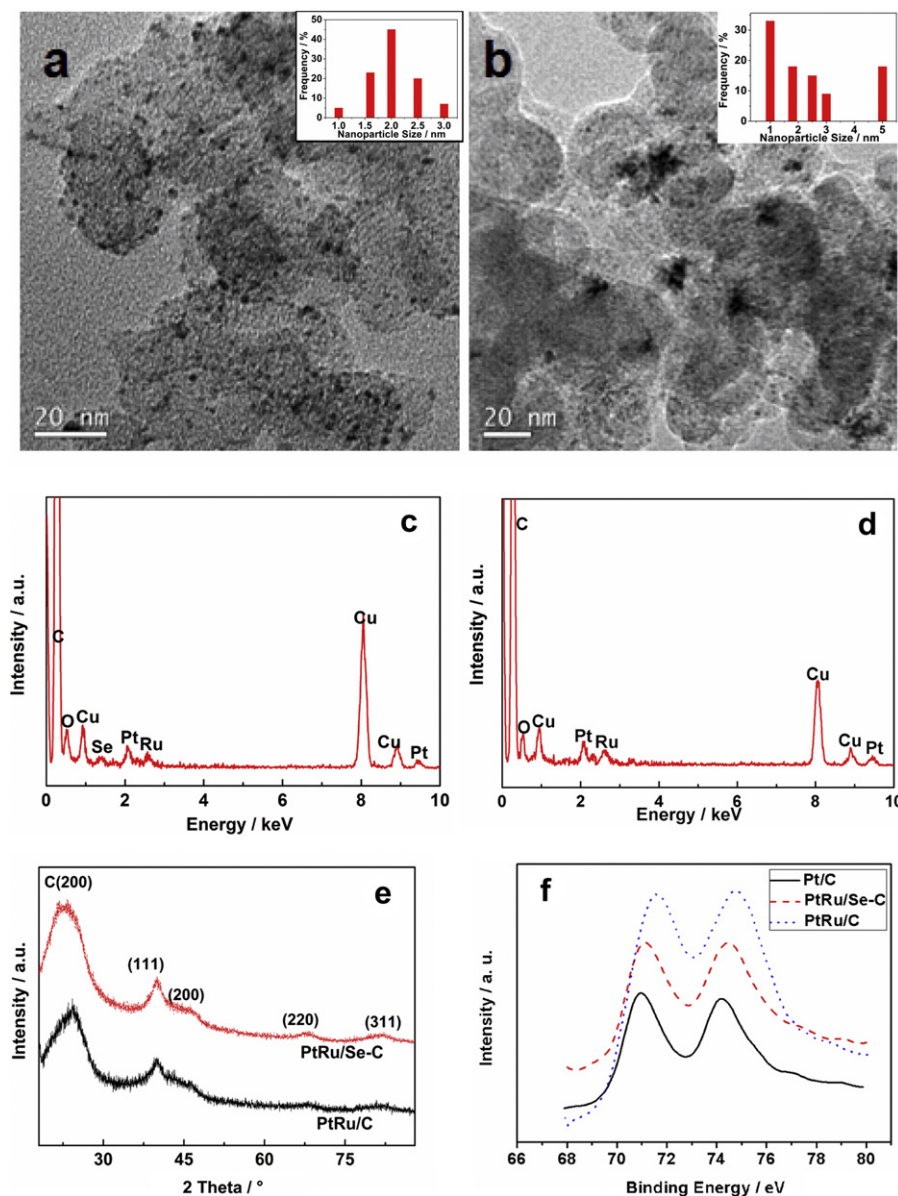


Fig. 3. (a) TEM image of the as-prepared PtRu/Se–C and (b) PtRu/C; Inserts in a and b: (top right corner) particle size distribution histograms of catalysts; (c) EDX of the as-prepared PtRu/Se–C and (d) PtRu/C; (e) XRD patterns of the PtRu/Se–C and PtRu/C samples; (f) XPS spectra for Pt 4f of Pt/C, PtRu/Se–C and PtRu/C.

From the hydrogen desorption peak areas, the electrochemical surface area (*ECSA*) of the PtRu/Se–C and PtRu/C catalysts were calculated from the CV curves according to literature methods [29]. The *ECSA* value of the PtRu/Se–C and PtRu/C catalysts is 84.6 and 69.5 m² g^{−1} of Pt, respectively. The higher *ECSA* of PtRu/Se–C catalyst compared to that of PtRu/C catalysts is due to much better dispersion of PtRu nanoparticles on the Se–C support.

Fig. 4b shows the CVs recorded for the PtRu/Se–C and PtRu/C catalysts in 0.5 M H₂SO₄ solution with 0.5 M CH₃OH in the potential range −0.2 to 1.0 V. Methanol oxidation occurs at a relatively lower potential and the current peak normalized to the *ECSA* is significantly higher for the PtRu/Se–C catalyst compared to that of the PtRu/C catalyst. The negative shift, ca. 52 mV, of the onset potential for the PtRu/Se–C catalyst (0.14 V vs. Ag/AgCl) compared with 0.19 V of the PtRu/C catalyst indicates good electrocatalytic activity of the PtRu/Se–C catalyst. It is also noted that the forward peak potential of methanol oxidation on the PtRu/Se–C catalyst is clearly more negative than that on PtRu/C catalysts. A specific activity for the methanol oxidation reaction for the PtRu/Se–C catalyst, 0.58 mA cm^{−2}, is

exhibited, while that of the PtRu/C catalyst is less at 0.44 mA cm^{−2}. The higher electrocatalytic activity of the PtRu/Se–C catalyst compared with the PtRu/C catalyst could be due to better dispersion.

Further, to compare the possible poisoning of the PtRu/Se–C and PtRu/C catalysts under short time continuous operation toward the methanol oxidation reaction, chronoamperometry tests, shown in Fig. 4c in 0.5 M H₂SO₄ solution containing methanol for 1000 s, were carried at a constant potential of 0.6 V vs. Ag/AgCl. In the initial time period the potentiostatic current density decreases rapidly for all the PtRu catalysts, which could be due to the formation intermediate species, such as CO_{ads}, CH₃OH_{ads}, and CHO_{ads}, during the methanol oxidation reaction [30]. At longer time periods although the current gradually decays for both catalysts, the PtRu/Se–C catalysts maintain a higher current density than the PtRu/C catalyst throughout all ranges up to 1000 s. The much higher tolerance of the PtRu/Se–C catalyst can be attributed to the change of electronic structure induced by the presence of Se.

First and second CVs during the stripping of a CO monolayer for PtRu/Se–C and PtRu/C catalysts are given in Fig. 4d. The adsorption

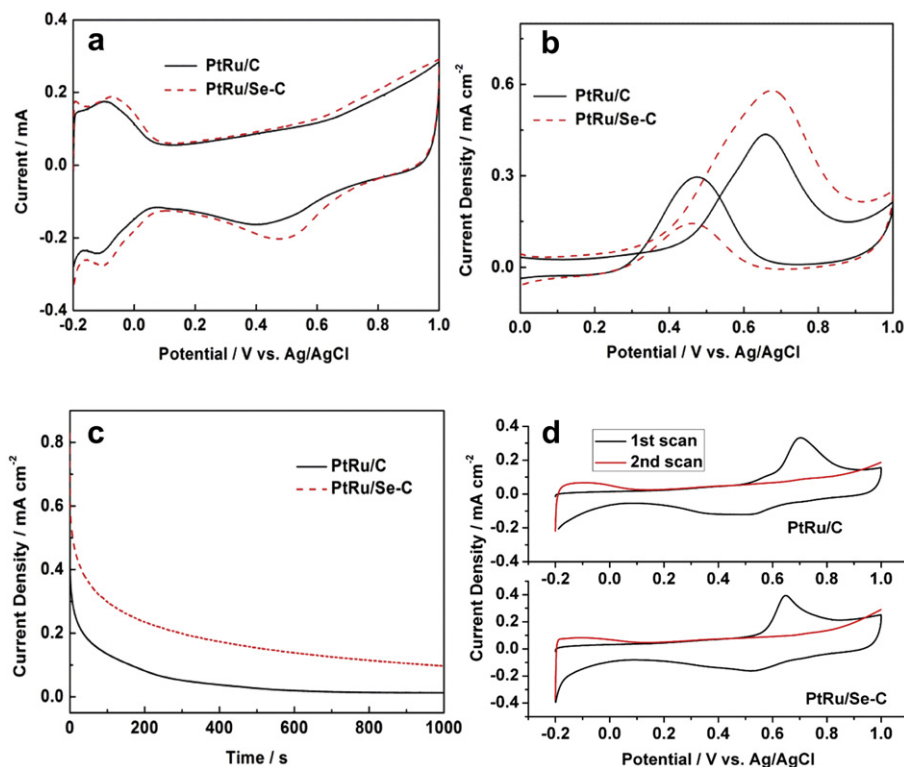


Fig. 4. Cyclic voltammograms of PtRu/Se–C and PtRu/C catalysts in (a) nitrogen saturated 0.5 M H₂SO₄; (b) nitrogen saturated 0.5 M H₂SO₄ + 0.5 M CH₃OH at a scan rate of 50 mV s⁻¹; and (c) chronoamperometry tests of methanol oxidation of PtRu/Se–C and PtRu/C catalysts in 0.5 M CH₃OH and 0.5 M H₂SO₄ solution; and (d) CO stripping curves of PtRu/Se–C and PtRu/C catalysts in 0.5 M H₂SO₄ solution at a scan rate of 50 mV s⁻¹.

time of CO was set to 5 min since a further increase in the adsorption time did not change the voltammograms [31]. A clean surface is recovered in the second cycle coinciding with the voltammogram described in the literature. For PtRu/C, the CO oxidation peak at 0.70 V attains the CO oxidation peak of PtRu/Se–C (0.65 V) shift to lower potential by ca. 50 mV. This result implies the PtRu/Se–C catalyst with the better CO tolerance than the PtRu/C catalyst, which could be attributed to the change of the Pt electronic structure induced by the presence of Se.

4. Conclusions

In summary, a new catalyst support, surface-selenium-enriched carbon, for the synthesis of noble metal catalysts, was successfully prepared by functionalized Vulcan carbon with non-covalently bonded Se atoms. PtRu NPs supported on the Se–C have better dispersion and higher electrochemical surface area than that on C without Se functionalization. Most importantly, the as-prepared PtRu/Se–C catalyst shows much better electrocatalytic activity toward methanol electrooxidation. The method reported here possesses a high potential for the fabrication of heteroatom functionalized carbon as a support for noble metal NPs in fuel cells.

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References

- [1] R.J.K. Wiltshire, C.R. King, A. Rose, P.P. Wells, H. Davies, M.P. Hogarth, D. Thompson, Brian Theobald, F.W. Mosselmans, M. Roberts, A.E. Russell, *Phys. Chem. Chem. Phys.* 11 (2009) 2305.
- [2] R. Ahmadi, M.K. Amini, *Int. J. Hydrogen Energy* 36 (2011) 7275.
- [3] G.S. Chai, S.B. Yoon, J.H. Kim, J.S. Yu, *Chem. Commun.* 23 (2004) 2766.
- [4] M. Kim, S. Hwang, J.S. Yu, *J. Mater. Chem.* 17 (2007) 1656.
- [5] G. Lee, J.H. Shim, H. Kang, K.M. Nam, H. Song, J.T. Park, *Chem. Commun.* 28 (2009) 5036.
- [6] P. Hernández-Fernández, R. Nuño, E. Fatás, J.L.G. Fierro, P. Ocón, *Int. J. Hydrogen Energy* 36 (2011) 8267.
- [7] K. Kardimi, T. Tsoufis, A. Tomou, B.J. Kooi, M.I. Prodromidis, D. Gournis, *Int. J. Hydrogen Energy* 37 (2012) 1243.
- [8] Y. Zhao, X. Yang, J. Tian, *Electrochim. Acta* 54 (2009) 7114.
- [9] B. Wu, D. Hu, Y. Kuang, Y. Yu, X. Zhang, J. Chen, *Chem. Commun.* 47 (2011) 5253.
- [10] A. Stein, Z. Wang, M.A. Fierke, *Adv. Mater.* 21 (2009) 265.
- [11] G. Wu, R. Swaidan, D. Li, N. Li, *Electrochim. Acta* 53 (2008) 7622.
- [12] R. Chetty, S. Kundu, W. Xia, M. Bron, W. Schuhmann, V. Chirila, W. Brandl, T. Reinecke, M. Muhler, *Electrochim. Acta* 54 (2009) 4208.
- [13] Z. Liu, F. Su, X. Zhang, S.W. Tay, *ACS Appl. Mater. Interfaces* 3 (2011) 3824.
- [14] Y. Zhou, K. Neyerlin, T.S. Olson, S. Pylypenko, J. Bult, H.N. Dinh, T. Gennett, Z. Shao, R. O'Hayre, *Energy Environ. Sci.* 3 (2010) 1437.
- [15] A. Abouimrane, D. Dambournet, K.W. Chapman, P.J. Chupas, W. Weng, K. Amine, *J. Am. Chem. Soc.* 134 (2012) 4505.
- [16] K.A. Kurak, A.B. Anderson, *J. Electrochem. Soc.* 157 (2010) B173.
- [17] Z. Wei, H. Guo, Z. Tang, *J. Power Sources* 62 (1996) 233.
- [18] P.K. Babu, A. Lewera, J.H. Chung, R. Hunger, W. Jaegermann, N. Alonso-Vante, A. Wieckowski, E. Oldfield, *J. Am. Chem. Soc.* 129 (2007) 15140.
- [19] R.F. Wang, S.J. Liao, H.Y. Liu, H. Meng, *J. Power Sources* 171 (2007) 471.
- [20] H. Li, D. Kang, H. Wang, R. Wang, *Int. J. Electrochem. Sci.* 6 (2011) 1058.
- [21] V. Georgakilas, A. Bourninos, D. Gournis, T. Tsoufis, C. Trapalis, A. Mateo-Alonso, M. Prato, *J. Am. Chem. Soc.* 130 (2008) 8733.
- [22] Y.J. Chen, G. Xiao, T.S. Wang, Q.Y. Ouyang, L.H. Qi, Y. Ma, P. Gao, C.L. Zhu, M.S. Cao, H.B. Jin, *J. Phys. Chem. C* 115 (2011) 13603.
- [23] C.S. Lee, S.E. Baker, M.S. Marcus, W. Yang, M.A. Eriksson, R.J. Hamers, *Nano Lett.* 4 (2004) 1713.
- [24] S. Kim, S.J. Park, *J. Power Sources* 159 (2006) 42.
- [25] F. Cataldo, *Polyhedron* 19 (2000) 681.
- [26] S.M.S. Kumar, J.S. Herrero, S. Irusta, K. Scott, *J. Electroanal. Chem.* 647 (2010) 211.
- [27] J. Inukai, D. Cao, A. Wieckowski, K.C. Chang, A. Menzel, V. Komanicky, H. You, *J. Phys. Chem. C* 111 (2007) 16889.
- [28] J. Prabhuram, T.S. Zhao, Z.X. Liang, R. Chen, *Electrochim. Acta* 52 (2007) 2649–2656.
- [29] H. Wang, R. Wang, H. Li, Q. Wang, J. Kang, Z. Lei, *Int. J. Hydrogen Energy* 36 (2011) 839.
- [30] J. Prabhuram, T.S. Zhao, Z.K. Tang, R. Chen, Z.X. Liang, *J. Phys. Chem. B* 110 (2006) 5245.
- [31] T.J. Schmidt, M. Noeske, H.A. Gasteiger, R.J. Behm, *Langmuir* 13 (1997) 2592.